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acid. A sample of acetic acid has the
following percentage composition: 39,9%
carbon 6,7% hydrogen 53,4% oxygen

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Determine the molecular formula of acetic acid if the molar mass of acetic acid is $60 \text{ g}\cdot\text{mol}^{-1}$. (6) 6.2.2 Give a reason why acetic acid is regarded as a monoprotic acid. (2)

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